



$V=1\text{L}$   
at  $25^\circ\text{C}$

$$m=7.41\text{g}$$

$$m=7.29\text{g}$$

$$M_{\text{Ca(OH)}_2}=74\text{g/mol}$$

$$M_{\text{HCl}}=36.5\text{g/mol}$$

$$n=0.1\text{mol}$$

$$n=0.2\text{mol}$$

0.1mol

$$c_{\text{CaCl}_2} = \frac{0.1\text{mol}}{1\text{L}} = 0.1\text{M}$$

$$\begin{array}{l} 1\text{ mol} - 2\text{ mol} \\ x - 0.2\text{ mol} \end{array} \quad x=0.1\text{mol}$$