

	<u>Run 1- Test</u>	<u>Run 2- Test</u>
KIO ₃ (g) added to 250 mL volumetric flask	0.41597	0.41597
Iodate Molar Mass (g/mol)	174.903	174.903
Iodate (moles) in 250 mL volumetric flask	0.00237829	0.00237829
KIO ₃ Injection (mL)	1	1
KIO ₃ Injection (L)	0.001	0.001
1N HCL injected into cell (mL)	1	1
10% KI solution injection (mL)	1	1
Iodate moles in titration cell	0.000010	0.000010
Iodine moles generated from reaction	0.000029	0.000029
Stannous Chloride Injection Volume (mL)	1	1
Stannous Chloride Injection Volume (L)	0.001	0.001
Sodium Thiosulfate Molarity (mol/L)	0.01	0.01
Sodium Thiosulfate Titre Volume (mL)	3.9683	4.3416
Sodium Thiosulfate Titre Volume (L)	0.0039683	0.0043416
Moles of sodium thiosulfate titrated	0.000039683	0.000043416
Excess Iodine titrated by Sodium Thiosulfate (moles)	1.98415E-05	0.000021708
Iodine moles that reacted with stannous	0.000009	0.000007
Iodate mole conversion from iodine	0.000003	0.000002
Potassium Iodate Moles Total	0.000003	0.000002
Potassium Iodate Molecular Weight	214.001	214.001
Potassium Iodate Mass Reacted with Stannous (g)	0.001	0.000
Potassium Iodate Density	3.89	3.89
Potassium Iodate Volume Reacted with Stannous (mL)	0.0024	0.0019
Conversion Factor Stannous (0.594 mg)	0.0014	0.0011
Stannous Chloride Concentration (ug/mL)	1.4337	1.1260
Mass of stannous chloride (mg) add to 100 mL volumetric flask	35	36.1
Theoretical Concentration		
µg SnCl ₂ (Anh.)/mL	294	303.24
10% Range Low	264.6	272.916
10% Range High	323.4	333.564