Mass KIO3 in 250 ml (g)	0.41597	0.41597
Molar mass of KIO3 (g/mol)	=126.9+48+39.1	=126.9+48+39.1
Moles KIO3 in 250 ml	=D2/D3	=E2/E3
ml KIO3 soln injected	1	1
moles KIO3 injected	=D6/250*D4	=E6/250*E4
ml 1N (?) HCl injected	1	1
moles HCl injected	=D8/1000*1	=E8/1000*1
mL 10% KI solution	1	1
moles KI injected	=D10/1000*100/166	=E10/1000*100/166
moles I2 generated	=D7*3	=E7*3
Molarity of Na2S2O3 sln	0.01	0.01
Titre volume ml	3.9683	4.3416
moles Na2S2O3 titrated	=D15/1000*D14	=E15/1000*E14
moles excess I2 titrated	=D16/2	=E16/2
moles I2 that reacted with Sn2+	=D12-D17	=E12-E17
moles of Sn2+ reacted	=D18	=E18
ml SnCl2 soln injected	1	1
molarity of SnCl2 soln	=D19/D20*1000	=E19/E20*1000
Molar mass of SnCl2	189.7	189.7
Concentration of SnCl2 (g/l)	=D21*D22	=E21*E22
Concentration of SnCl2 (µg/ml)	=D23*1000	=E23*1000